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Instead, the quantum-
mechanical model for
the atom gives the
probabilities of finding

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an electron in a particular region around the nucleus. In this experiment, we model the probable locations of electrons around the nucleus. We will use dried beans or peas to represent the electrons to help us visualize the regions of low and high density.

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a given moment
cannot be predicted.

The region of space in

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which the electron can probably be found is called an s orbital.

Orbitals are often referred to as

□electron clouds□

because they are not absolute. In fact, there are times when an electron may be found outside the orbital.

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purpose of this lab is

to investigate the

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According to modern
atomic theory, we
cannot be sure of the
exact location of
electrons in an atom.

We predict that
electrons will be

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relatively close to the nucleus (because the electrons are negatively charged and the nucleus is positively charged).

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$N=3$, $L=1$, $M=1$,

$M=1/2$. This electron

is found within the 2p

orbital, which has the

dumbbell shape,

primarily based on the

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z axis. This electron would share this orbital with another electron, but would have a negative spin.

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The position of an
electron in an atom at
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cannot be predicted.

The region of space in which the electron can probably be found is called an orbital.

Orbitals are often referred to as

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because they are not absolute. In fact, there are times when an electron may be found outside the orbital.

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Introduce electrons in atoms by dropping darts! Lab provides a discussion of probability and electron density.

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(d) If we count the
number of dots in

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each spherical shell, we obtain the total probability of finding the electron at a given value of r . Because the surface area of each shell increases more rapidly with increasing r than the electron probability density decreases, a plot of electron probability versus r (the radial probability)

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shows a peak. This peak corresponds to the most probable radius for the electron, 52.9 pm, which is exactly the radius predicted by Bohr's model of the hydrogen ...

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